Limiting and excess reactants worksheet doc

Continue

Limiting Reactant

Name ______________________

Date ____________ Period ______

Show ALL of your work for credit on this assignment!

1. _____ Cu(s) + _____ $AgNO_3$(aq) → _____ $Cu(NO_3)_2$(aq) + _____ Ag(s)

   If 2.5 moles of copper and 5.5 moles of silver nitrate are available to react, what is the limiting reactant?

2. _____ CaO(s) + _____ H_2O(l) → _____ $Ca(OH)_2$(aq)

   How many grams of calcium hydroxide will be formed in this reaction when 4.44 g of calcium oxide and 7.77 g of water are available to react? Also identify the limiting and excess reactants.

Limiting Reagent Worksheet #1

1. Given the following reaction: (Balance the equation first!)

C_3H_8 + O_2 ——> CO_2 + H_2O

a) If you start with 14.8 g of C_3H_8 and 3.44 g of O_2, determine the limiting reagent

b) determine the number of moles of carbon dioxide produced

c) determine the number of grams of H_2O produced

d) determine the number of grams of excess reagent left

2. Given the following equation:

$Al_2(SO_3)_3$ + 6 NaOH ——> 3 Na_2SO_3 + 2 $Al(OH)_3$

a) If 10.0 g of $Al_2(SO_3)_3$ is reacted with 10.0 g of NaOH, determine the limiting reagent

b) Determine the number of moles of $Al(OH)_3$ produced

c) Determine the number of grams of Na_2SO_3 produced

d) Determine the number of grams of excess reagent left over in the reaction

3. Given the following equation:

Al_2O_3 + Fe ——> Fe_3O_4 + Al

a) If 25.4 g of Al_2O_3 is reacted with 10.2 g of Fe, determine the limiting reagent

b) Determine the number of moles of Al produced

c) Determine the number of grams of Fe_3O_4 produced

d) Determine the number of grams of excess reagent left over in the reaction

Name: ______________________ **Date:** _________ **Period:** ___

Chemistry Worksheet: Limiting Reactant Worksheet

1. Consider the following reaction: 2 Al + 6 HBr → 2 $AlBr_3$ + 3 H_2

a. When 3.22 moles of Al reacts with 4.96 moles of HBr, what is the limiting reactant?

b. how many moles of H_2 are formed?

c. For the reactant in excess, how many moles are left over at the end of the reaction?

2. Consider the following reaction: 3 Si + 2 N_2 → Si_3N_4

a. When 21.44 moles of Si reacts with 17.62 moles of N_2, what is the limiting reactant?

b. how many moles of Si_3N_4 are formed?

c. For the reactant in excess, how many moles are left over at the end of the reaction?

3. Consider the following reaction: 2 $CuCl_2$ + 4 KI → 2 CuI + 4 KCl + I_2

a. When 0.56 moles of $CuCl_2$ reacts with 0.64 moles of KI, what is the limiting reactant?

b. how many moles of I_2 are formed?

c. For the reactant in excess, how many moles are left over at the end of the reaction?

4. Consider the following reaction: 4 FeS_2 + 11 O_2 → 2 Fe_2O_3 + 8 SO_2

a. When 26.62 moles of FeS_2 reacts with 5.44 moles of O_2, what is the limiting reactant?

b. how many moles of SO_2 are formed?

c. For the reactant in excess, how many moles are left over at the end of the reaction?

Honors Chemistry 1B
Limit Reactant and Percent Yield Worksheet
(with excess calculation)

Name: ______________________

STO.4	Solve stoichiometric problems from a balanced chemical equation.	3
STO.5	Differentiate between the actual yield and theoretical yield of a chemical reaction.	3
STO.6	Calculate the theoretical yield from a balanced chemical equation using dimensional analysis.	3
STO.7	Determine the limiting reactant from stoichiometric calculations.	3

1. Balance the equation for the reaction given below:

   $CuCl_2$ + $NaNO_3$ → $Cu(NO_3)_2$ + NaCl

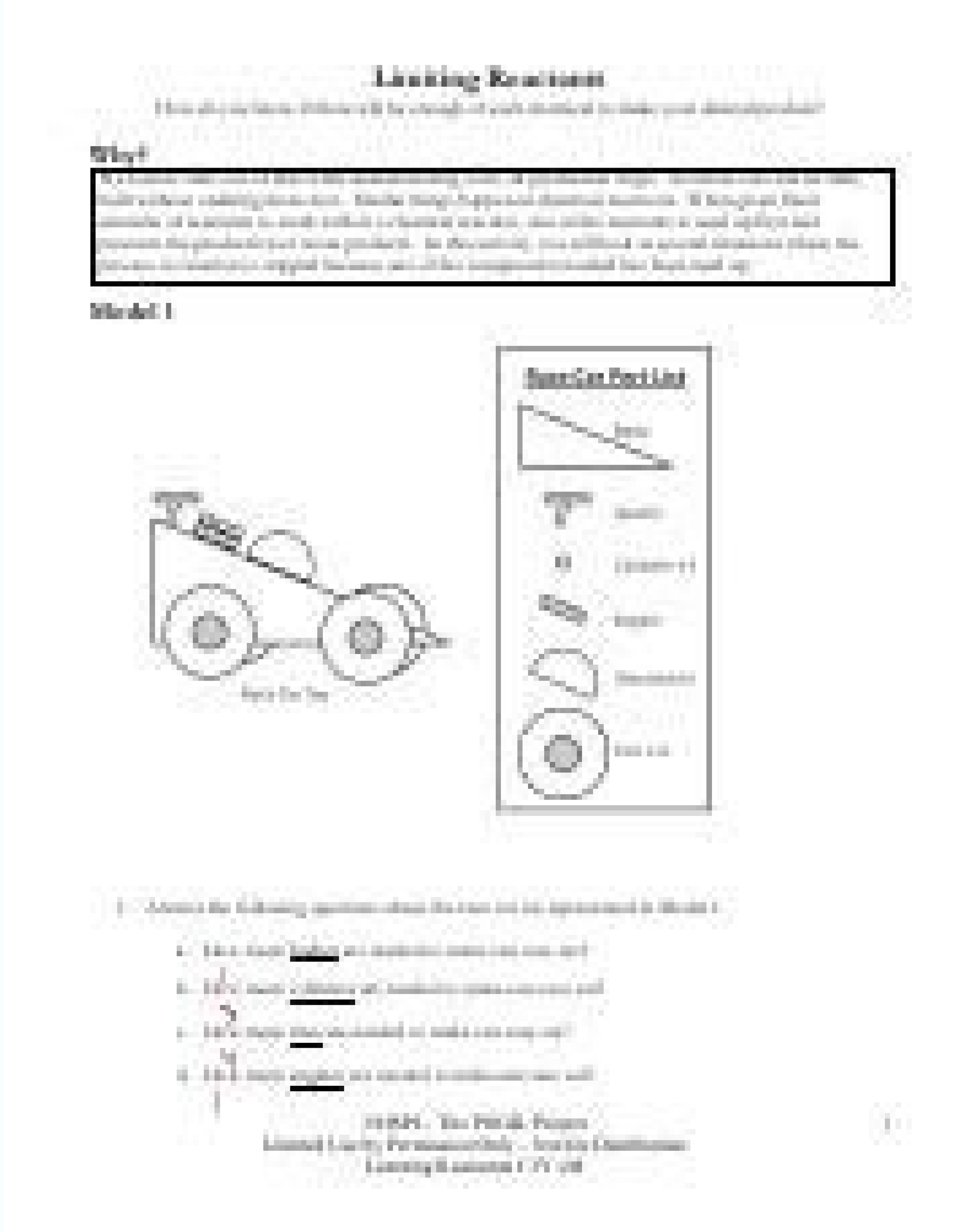
   a) If 15 grams of copper (II) chloride react with 20. grams of sodium nitrate, how much sodium chloride can be formed?

   b) What is the name of the limiting reagent? ______________

   c) How much of the excess reagent is left over in this reaction?

   d) If 11.3 grams of sodium chloride are formed in the reaction, what is the percent yield of this reaction?

Modified from - Limiting Reactant and Percent Yield Wkst.pdf
Blake – 3/2015



Limiting and excess reactants worksheet pdf. Limiting and excess reactants worksheet answers pdf. Limiting and excess reactants worksheet with answers.

Klaus Grawe hat mit seiner Konsistenztheorie (2000, 2004) versucht, das psychische Funktionieren des Menschen vor dem Hintergrund der Befriedigung psychologischer Grundbedürfnisse, zu erklären. Er bemühte sich hierbei, seine Theorie erfahrungswissenschaftlich zu untermauern. Grawe zufolge streben alle Organismen nach Konsistenz, d.h. nach einer Passung zwischen den inneren Bedürfnissen (psychische Prozesse) und dem Erleben in der Realität (neuronale Prozesse). Je höher die Konsistenz ist, desto gesünder ist der Organismus. Neben dem Streben nach Konsistenz, welches auf der Systemebene abläuft, unterscheidet Grawe drei weitere Ebenen: Die erste Ebene bilden hierbei die Grundbedürfnisse, deren Befriedigung die größte Wichtigkeit hat. Grawe geht davon aus, dass es insgesamt vier Grundbedürfnisse gibt: Bedürfnis nach Orientierung und Kontrolle Bedürfnis nach Lustgewinn und Unlustvermeidung Bedürfnis nach Bindung Bedürfnis nach Selbstwertschutz und Selbstwerterhöhung. Die Grundbedürfnisse sind universell bei jedem Menschen zu finden. Die Methoden und Wege, diese Grundbedürfnisse zu befriedigen, sind jedoch von Mensch zu Mensch unterschiedlich. Sie sind u.a. abhängig von seinen Sozialisationserfahrungen. Auf der nächsten Ebene finden sich daher die sogenannten motivationalen Schemata, welche sich im Laufe der Entwicklung, in Interaktion mit der Umwelt, bei allen Menschen entwickeln. Diese lassen sich unterteilen in: Intentionale Schemata (Annäherungsziele) und Vermeidungsschemata (Vermeidungsziele) Die Annäherungsziele dienen dazu, die Grundbedürfnisse der Person zu befriedigen. Die Vermeidungsziele hingegen dienen dem Schutz vor Bedrohung, Verletzung oder Frustration der Grundbedürfnisse. In der Regel sind beide Systeme gleichzeitig aktiv und der Organismus strebt nach einer optimalen Bilanz zwischen allen aktivierten motivationalen Zielen. Die einzelnen Schemata können jedoch auch sehr unterschiedlich stark ausgeprägt sein bei verschiedenen Menschen, je nachdem, welche prägenden Erfahrungen sie gemacht haben. So gibt es Menschen mit besonders stark entwickelten Vermeidungszielen und wieder andere, deren Annäherungsziele überwiegen. Diskordanz liegt dann vor, wenn Schemata bzw. Ziele untereinander in Konflikt geraten, d.h. wenn z.B. Annäherungs- und Vermeidungsziele gleichzeitig aktiviert werden und sich dadurch gegenseitig hemmen. Die motivationalen Konflikte können jedoch auch untereinander (Annäherungs-/ Annäherungskonflikt, Vermeidungs-/ Vermeidungskonflikt) bestehen. Auf der dritten Ebene werden die motivationalen Ziele mit den realen Wahrnehmungen abgeglichen (Ebene des Erlebens und Verhaltens). Stimmen die Annäherungsziele und die realen Erfahrungen nicht überein, dann entsteht Inkongruenz. Das heißt, die motivationalen Ziele werden nicht erreicht. Hieraus resultieren u.a. negative Emotionen. Konsistenz entsteht dann, wenn Grundbedürfnisse ausgeglichen und motivationale Ziele erreicht werden. Ist das Ziel (Konsistenz) erreicht, kommt es zu positiven Emotionen. Inkonsistenz ergibt sich aus Inkongruenz und Diskordanz und gilt als wichtiger Faktor bei der Entstehung psychischer Störungen. Therapeutische Relevanz: Mit Hilfe der vertikalen Verhaltensanalyse oder Plananalyse können zu Beginn der Therapie Annäherungs- und Vermeidungsziele herausgearbeitet werden und hierdurch Inkonsistenzen und motivationale Konflikte aufgedeckt werden. Grawe geht zudem davon aus, dass eine Therapie dann wirkungsvoll ist, wenn die Konsistenzerfahrung gesteigert wird. Dies geschieht auf zwei Arten: Indem Inkongruenzquellen reduziert werden (d.h. zum Beispiel, motivationale Konflikte aufgelöst werden) Und es zu bedürfnisbefriedigenden Erfahrungen kommt (d.h. Bedürfnisbefriedigung unter anderem durch komplementäre Beziehungsgestaltung). Die Therapie nach der Konsistenztheorie wird auch als Neuropsychotherapie bezeichnet, weil sich die Veränderungsprozesse auch auf der neuronalen Ebene im Gehirn abzeichnen. Sichtbar gemacht werden kann dies u.a. mit Hilfe der Methode des Neurofeedbacks. Quellen: Grawe, K. (2000). Psychologische Therapie. Göttingen: Hogrefe-Verlag. Grosse Holtforth, M. and K. Grawe (2004). Konfliktdiagnostik aus der Perspektive der Konsistenztheorie. Lernen an der Praxis. OPD und Qualitätssicherung in der Psychodynamischen Psychotherapie. R. W. Dahlenbender, P. Buchheim and G. Schüssler. Bern, Huber. Grosse Holtforth, M. and K. Grawe (2004). "Inkongruenz und Fallkonzeption in der Psychologischen Therapie." Verhaltenstherapie und psychosoziale Praxis 36(1): 921. Kandale, O. & Rugenstein, K. (2016). Das Repetitorium - Lehr- und Lernbuch für die schriftlichen Abschlussprüfungen zum Psychologischen Psychotherapeuten und zum Kinder- und Jugendlichenpsychotherapeuten. Berlin: Deutscher Psychologen Verlag GmbH. Assignment Essay Help Our professional team of writers ensures top-quality custom essay writing services. We strive to ensure that every paper is crafted with getting you the highest grade in mind. Best Customer Support Service Get 24⁄7 customer support help when you place a homework help service order with us. We will guide you on how to place your essay help, proofreading and editing your draft – fixing the grammar, spelling, or formatting of your paper easily and cheaply. Affordable Essay Writing Service We guarantee a perfect price-quality balance to all students. The more pages you order, the less you pay. We can also offer you a custom pricing if you feel that our pricing doesn't really feel meet your needs. Full PDF PackageDownload Full PDF PackageThis PaperA short summary of this paper22 Full PDFs related to this paperDownloadPDF Pack DP Chemistry Review Topic 1: Quantitative chemistry 1.1 The mole concept and Avogadro s constant Assessment statement Apply the mole concept to substances. Determine the number of particles and the amount More information CHAPTER 3: MATTER Active Learning Questions: 1-6, 9, 13-14; End-of-Chapter Questions: 1-18, 20, 24-32, 38-42, 44, 49-52, 55-56, 61-64 3.1 MATTER Matter: Anything that has mass and occupies volume We study More information Experiment 8 - Double Displacement Reactions A double displacement reaction involves two ionic compounds that are dissolved in water. In a double displacement reaction, it appears as though the ions are More information 1 P age Module 5: Combustion Technology Lecture 33: Combustion air calculation 2 P age Keywords: Heat of combustion, stoichiometric air, excess air, natural gas combustion Combustion air calculation The More information Cautions Butane is toxic and flammable. No OPEN Flames should be used in this experiment. Purpose The purpose of this experiment is to determine the molar mass of butane using Dalton s Law of Partial Pressures More information Basics of Kraft Pulping & Recovery Process Art J. Ragauskas Institute of Paper Science and Technology Georgia Institute of Technology Outline History Goals Process Overview Kraft Pulping Process Kraft More information Page 1 of 76 1.0 PURPOSE The purpose of the Wastewater Treatment System is to remove contaminates from plant wastewater so that it may be sent to the Final Plant Effluent Tank and eventually discharged More information Chemistry Form 3 Page 62 Ms. R. Buttigieg Unit 6 The Mole Concept See Chemistry for You Chapter 28 pg. 352-363 See GCSE Chemistry Chapter 5 pg. 70-79 6.1 Relative atomic mass. The relative atomic mass More information Cambridge International Examinations Cambridge International General Certificate of Secondary Education *0123456789* CHEMISTRY 0620/03 Paper 3 Theory (Core) For Examination from 2016 SPECIMEN PAPER 1 hour More information DETERMINING THE ENTHALPY OF FORMATION OF CaCO 3 Standard Enthalpy Change Standard Enthalpy Change for a reaction, symbolized as H 0 298, is defined as The enthalpy change when the molar quantities of reactants More information 1. The average kinetic energy of water molecules increases when 1) H 2 O(s) changes to H 2 O( ) at 0ºC 3) H 2 O( ) at 10ºC changes to H 2 O( ) at 20ºC 2) H 2 O( ) changes to H 2 O(s) at 0ºC 4) H 2 O( ) More information 1 Experiment 7: Titration of an Antacid Objective: In this experiment, you will standardize a solution of base using the analytical technique known as titration. Using this standardized solution, you will More information The Mole Atomic mass units and atoms are not convenient units to work with. The concept of the mole was invented. This was the number of atoms of carbon-12 that were needed to make 12 g of carbon. 1 mole More information EXPERIMENT 12: Empirical Formula of a Compound INTRODUCTION Chemical formulas indicate the composition of compounds. A formula that gives only the simplest ratio of the relative number of atoms in a compound More information Helsinki University of Technology Department of Mechanical Engineering Energy Engineering and Environmental Protection Publications Steam Boiler Technology ebook Espoo 2002 Boiler Calculations Sebastian More information 1 P age Module 5: Combustion Technology Lecture 34: Calculation of calorific value of fuels 2 P age Keywords : Gross calorific value, Net calorific value, enthalpy change, bomb calorimeter 5.3 Calculation More information All of the chemical changes you observed in the last Investigation were the result of chemical reactions. A chemical reaction involves a rearrangement of atoms in one or more reactants to form one or more More information GCSE CHEMISTRY Higher Tier Chemistry 1H H Specimen 2018 Time allowed: 1 hour 45 minutes Materials For this paper you must have: a ruler a calculator the periodic table (enclosed). Instructions Answer all More information 11-1 Stoichiometry What is stoichiometry? Calculations that relate the quantities of substances. It is the study of quantitative (measurable amounts) relationships in chemical reactions and equations. More information Honors Chemistry: Unit 6 Test Stoichiometry PRACTICE TEST ANSWER KEY Page 1 1. 2. 3. 4. 5. 6. Question What is a symbolic representation of a chemical reaction? What 3 things (values) is a mole of a chemical More information Mass, Moles, & Molar Mass Relative quantities of isotopes in a natural occurring element (%) E.g. Carbon has 2 isotopes C-12 and C-13. Of Carbon s two isotopes, there is 98.9% C-12 and 11.1% C-13. Find More information Chemistry 13: States of Matter Name: Period: Date: Chemistry Content Standard: Gases and Their Properties The kinetic molecular theory describes the motion of atoms and molecules and explains the properties More information Chapter 5 Student Reading THE POLARITY OF THE WATER MOLECULE Wonderful water Water is an amazing substance. We drink it, cook and wash with it, swim and play in it, and use it for lots of other purposes. More information Chapter 8 - Chemical Equations and Reactions 8-1 Describing Chemical Reactions I. Introduction A. Reactants 1. Original substances entering into a chemical rxn B. Products 1. The resulting substances from More information CHATER 3. The atmosphere is a homogeneous mixture (a solution) of gases.. Solids and liquids have essentially fixed volumes and are not able to be compressed easily. have volumes that depend on their conditions, More information IBP 2778_10 HIGH EFFICIENCY ON CO2 REMOVAL IN NATURAL GAS WITH UCARSOL SOLVENTS Thiago V. Alonso 1 Copyright 2010, Brazilian Petroleum, Gas and Biofuels Institute - IBP This Technical Paper was prepared More information Experiment 3 Separation by Solvent Extraction Objectives To separate a mixture consisting of a carboxylic acid and a neutral compound by using solvent extraction techniques. Introduction Frequently, organic More information Chapter 2.2: Boilers Part I: Objective type Questions and Answers 1. The minimum capacity of any closed vessel which generates steam under Indian Boilers Regulation Act is. a) 2.275 liters b) 22.75 kilo More information Page : 1 of 38 Project Engineering Standard www.klmtechgroup.com KLM Technology #03-12 Block Aronia, Jalan Sri Perkasa 2 Taman Tampoi Utama 81200 Johor Bahru Malaysia CAUSTIC AND CHEMICAL TABLE OF CONTENT More information Lab 22 Properties of Acids and Bases TN Standard 4.2: The student will investigate the characteristics of acids and bases. Have you ever brushed your teeth and then drank a glass of orange juice? What More information Balancing Chemical Equations A Chemical Equation: is a representation of a chemical reaction in terms of chemical formulas Example: 1. Word Description of a Chemical Reaction When methane gas (CH 4) burns More information Lecture 35: Atmosphere in Furnaces Contents: Selection of atmosphere: Gases and their behavior: Prepared atmospheres Protective atmospheres applications Atmosphere volume requirements Atmosphere sensors More information DIN 2403 Identification of pipelines according to the fluid conveyed. Marking of pipes according to fluid transported 1 Field of application This standard specifies the colours for the identification of More information The Single Absorption Scrubbing Sulfuric Acid Process Leonard J. Friedman, Samantha J. Friedman Acid Engineering & Consulting, Inc. 17770 Deauville Lane, 33496, USA LJ.Friedman@Acideng.com Keywords: Sulfuric More information Stoichiometry Lecture Examples Answer Key Ex. 1 Balance the following chemical equations: 3 NaBr + 1 H 3 PO 4 3 HBr + 1 Na 3 PO 4 2 C 3 H 5 N 3 O 9 6 CO 2 + 3 N 2 + 5 H 2 O + 9 O 2 2 Ca(OH) 2 + 2 SO 2 More information By Philip Sutter Pick Heaters, Inc. DIRECT STEAM INJECTION HOT WATER SYSTEMS FOR JACKETED HEATING INTRODUCTION Many process plants currently use steam or hot water to heat jacketed devices such as tanks, More information AMMONIA AND UREA PRODUCTION Urea (NH 2 CONH 2) is of great importance to the agriculture industry as a nitrogen-rich fertiliser. In Kapuni, Petrochem manufacture ammonia and then convert the majority More information Chapter 3 ATOMS AND MOLECULES Multiple Choice Questions 1. Which of the following correctly represents 360 g of water? (i) 2 moles of H 2 0 (ii) 20 moles of water (iii) 6.022 10 23 molecules of water (iv) More information GS106 Chemical Bonds and Chemistry of Water c:wou:gs106:sp2002:chem.wpd I. Introduction A. Hierarchy of chemical substances 1. atoms of elements - smallest particles of matter with unique physical and More information CHEM 15 HOUR EXAM III 28-OCT-99 NAME (please print) 1. a. given: Ni (s) + 4 CO (g) = Ni(CO) 4 (g) H Rxn = -163 k/mole determine H f for Ni(CO) 4 (g) b. given: Cr (s) + 6 CO (g) = Cr(CO) 6 (g) H Rxn = -26 More information Chapter 8: Chemical Equations and Reactions I. Describing Chemical Reactions A. A chemical reaction is the process by which one or more substances are changed into one or more different substances. A chemical More information Production of R-134a Background In the 1930 s, chlorofluorocarbons (CFC s) were developed as a supposedly safe alternative to ammonia and sulfur dioxide refrigerants. While sulfur dioxide is toxic and More information Chapter 3 Formulas, Equations and Moles Interpreting Chemical Equations You can interpret a balanced chemical equation in many ways. On a microscopic level, two molecules of H 2 react with one molecule More information QUESTION (2012:3) (i) Complete the table below showing the conjugate acids and bases. Conjugate acid Conjugate base - HCO 3 2 CO 3 H 2 O OH HCN CN - (ii) HPO 4 2 (aq) Write equations for the reactions More information CHEMICAL REACTIONS OF COPPER AND PERCENT YIELD Objective To gain familiarity with basic laboratory procedures, some chemistry of a typical transition element, and the concept of percent yield. Apparatus More information ACIDS AND BASES Mild acids and bases are used in cooking (their reaction makes biscuits and bread rise). Acids such as those in our stomachs eat away at food or digest it. Strong acids and bases are used More information The position of hydrogen in the reactivity series Hydrogen, although not a metal, is included in the reactivity series because it, like metals, can be displaced from aqueous solution, only this time the More information Balancing Act Teacher Information Objectives In this activity, students neutralize a base with an acid. Students determine the point of neutralization of an acid mixed with a base while they: Recognize More information Chemistry B11 Chapter 4 Chemical reactions Chemical reactions are classified into five groups: A + B AB Synthesis reactions (Combination) H + O H O AB A + B Decomposition reactions (Analysis) NaCl Na +Cl More information Reminder: These notes are meant to supplement, not replace, the laboratory manual. Oxidation of Cyclohexanol to Cyclohexanone History and Application: Oxidation reactions are incredibly important in the More information 1 Experiment 3: Extraction: Separation of an Acidic, a Basic and a Neutral Substance Read pp 142-155, 161-162, Chapter 10 and pp 163-173, Chapter 11, in LTOC. View the videos: 4.2 Extraction (Macroscale); More information COMBUSTION In order to operate a heat engine we need a hot source together with a cold sink Occasionally these occur together in nature eg:- geothermal sites or solar powered engines, but usually the heat More information Skills Worksheet Problem Solving Stoichiometry of Gases Now that you have worked with relationships among moles, mass, and volumes of gases, you can easily put these to work in stoichiometry calculations. More information Notes Chapter 9 Limiting Reagent Sample Problems Page 1 Problem 1: Sodium chloride can be prepared by the reaction of sodium metal with chlorine gas. Suppose that 6.70 Na reacts with 3.20 Cl 2. A. What More information #3. Acid - Base Titrations 27 EXPERIMENT 3. ACID-BASE TITRATIONS: DETERMINATION OF CARBONATE BY TITRATION WITH HYDROCHLORIC ACID BACKGROUND Carbonate Equilibria In this experiment a solution of hydrochloric More information 5.2. Vaporizers - Types and Usage 5.2.1. General Vaporizers are constructed in numerous designs and operated in many modes. Depending upon the service application the design, construction, inspection, More information Types of Reactions Objectives The objectives of this laboratory are as follows: To perform and observe the results of a variety of chemical reactions. To become familiar with the observable signs of chemical More information Moles 1 MOLES The mole the standard unit of amount of a substance (mol) the number of particles in a mole is known as Avogadro s constant (N A) Avogadro s constant has a value of 6.02 x 10 23 mol -1. More information Helsinki University of Technology Department of Mechanical Engineering Energy Engineering and Environmental Protection Publications Steam Boiler Technology ebook Espoo 2002 Basics of Steam Generation Sebastian More information Decaffeination of Raw, Green offee Beans Using Supercritical O 2 Background The worldwide coffee market and the industry that supplies it are among the world s largest. offee as a world commodity is second More information Chapter 2 Chemical and Physical Properties of Sulphur Dioxide and Sulphur Trioxide 2.1 Introduction In order to appreciate the impact of the properties of liquid sulphur dioxide and liquid sulphur trioxide More information Waste Incineration Plants Modern Technology for a better Environmental Welcome at Hafner! We Manufacture Systems for Energy Recovery from Wastes and Biomass as well as for Treatment of Hazardous Wastes. More information Chemical versus Physical Changes Permission to Copy - This document may be reproduced for non-commercial educational purposes Copyright 2009 General Electric Company What are physical and chemical changes? More information Chemistry R: Form TR9.13A TEST 9 REVIEW Name Date Period Test Review # 9 Collision theory. In order for a reaction to occur, particles of the reactant must collide. Not all collisions cause reactions. More information Chapter 14 Solutions 1 14.1 General properties of solutions solution a system in which one or more substances are homogeneously mixed or dissolved in another substance two components in a solution: solute More information Chemical Reactions in Water Ron Robertson r2 f:\files\courses\1110-20\2010 possible slides for web\waterchemtrans.doc Properties of Compounds in Water Electrolytes and nonelectrolytes Water soluble compounds More information Centre Number 71 Candidate Number ADVANCED SUBSIDIARY (AS) General Certificate of Education January 2011 Chemistry Assessment Unit AS 1 assessing Basic Concepts in Physical and Inorganic Chemistry [AC111] More information Chemical Equations & Stoichiometry Chapter Goals Balance equations for simple chemical reactions. Perform stoichiometry calculations using balanced chemical equations. Understand the meaning of the term More information Physical Changes and Chemical Reactions Gezahegn Chaka, Ph.D., and Sudha Madhugiri, Ph.D., Collin College Department of Chemistry Objectives Introduction To observe physical and chemical changes. To identify More information Chapter 10 MULTIPLE CHOICE. Choose the one alternative that best completes the statement or answers the question. 1) A gas at a pressure of 10.0 Pa exerts a force of N on an area of 5.5 m2. A) 1.8 B) 0.55 More information THERMODYNAMICS TUTORIAL 5 HEAT PUMPS AND REFRIGERATION On completion of this tutorial you should be able to do the following. Discuss the merits of different refrigerants. Use thermodynamic tables for More information Cambridge International Examinations Cambridge International General Certificate of Secondary Education *0123456789* CHEMISTRY 0620/04 Paper 4 Theory (Extended) For Examination from 2016 SPECIMEN PAPER More information Molar Volume of Carbon Dioxide Reading assignment: Julia Burdge, Chemistry 3rd edition, Chapter 10. Goals To determine the molar volume of carbon dioxide gas and the amount of sodium carbonate in a sample. More information Stoichiometry 1 The total number of moles represented by 20 grams of calcium carbonate is (1) 1; (2) 2; (3) 01; (4) 02 2 A 44 gram sample of a hydrate was heated until the water of hydration was driven More information 3 Formulae, stoichiometry and the mole concept Content 3.1 Symbols, Formulae and Chemical equations 3.2 Concept of Relative Mass 3.3 Mole Concept and Stoichiometry Learning Outcomes Candidates should be More information Moles 1 MOLES The mole the standard unit of amount of a substance the number of particles in a mole is known as Avogadro s constant (L) Avogadro s constant has a value of 6.023 x 10 23 mol -1. Example More information SULFUR RECOVERY UNIT Thermal Oxidizer BURNERS FLARES INCINERATORS PARTS & SERVICE SULFUR RECOVERY UNIT Thermal Oxidizer Tail Gas Thermal Oxidizer designed and built to GOST-R requirements. Zeeco can meet More information Heating Water by Direct Steam Injection Producing hot water by direct steam injection provides a solution where large volumes of hot water at precise temperatures are required, and where energy and space More information Chapter 3. Stoichiometry: Mole-Mass Relationships in Chemical Reactions Concept 1. The meaning and usefulness of the mole The mole (or mol) represents a certain number of objects. SI def.: the amount of More information IGEMA BOILER LEVEL & TDS CONTROLS IGEMA offers boiler level and TDS control products of the highest quality standard, being certified to ISO 9001. Made in Germany, IGEMA products are manufactured in compliance More information Extraction: Separation of Acidic Substances Chemists frequently find it necessary to separate a mixture of compounds by moving a component from one solution or mixture to another. The process most often More information Thermo 1 (MEP 261) Thermodynamics An Engineering Approach Yunus A. Cengel & Michael A. Boles 7 th Edition, McGraw-Hill Companies, ISBN-978-0-07-352932-5, 2008 Sheet 5:Chapter 5 5 1C Name four physical More information Chemical Changes Name(s) In the particle model of matter, individual atoms can be bound tightly to other atoms to form molecules. For example, water molecules are made up of two hydrogen atoms bound to More information Sample Questions REVISED FIRST CLASS PARTS A1, A2, AND A3 (NOTE: these questions are intended as representations of the style of questions that may appear on examinations. They are not intended as study More information Balancing chemical reaction equations (stoichiometry) This worksheet and all related files are licensed under the Creative Commons Attribution License, version 1.0. To view a copy of this license, visit More information Week 9 Equilibria Involving Acids & Bases Acidic and basic solutions Self-ionisation of water Through reaction with itself: The concentration of water in aqueous solutions is virtually constant at about More information The role of CO 2 in pool water A series of e-mail articles from the research group onbalance, January 2006 The role of CO 2 in pool water #1 While some service techs go about their business taking care More information Lecture 28 Potash Fertilizers - Potassium Sulphate Potassium Sulfate Potassium sulfate is the second largest tonnage potassium compound and it is also used primarily as a fertilizer. The sulfate or other More information Chem. I Notes Ch. 12, part 2 Using Moles NOTE: Vocabulary terms are in boldface and underlined. Supporting details are in italics. 1 MOLE = 6.02 x 10 23 representative particles (representative particles More information Contents Getting the most from this book...4 About this book....5 Content Guidance Topic 1 Atomic structure and the periodic table...8 Topic 2 Bonding and structure...14 Topic 2A Bonding....14 Topic 2B More information Increasing Natural Gas Boiler Efficiency by Capturing Waste Energy from Flue Gas Mark Schiffhauer, ATSI Engineering Services Cameron Veitch, Combustion and Energy Systems Scott Larsen, New York State Energy More information Continuous flow direct water heating for potable hot water An independently produced White Paper for Rinnai UK 2013 www.rinnaiuk.com In the 35 years since direct hot water systems entered the UK commercial More information THE HUMIDITY/MOISTURE HANDBOOK Table of Contents Introduction... 3 Relative Humidity... 3 Partial Pressure... 4 Saturation Pressure (Ps)... 5 Other Absolute Moisture Scales... 8 % Moisture by Volume (%M More information 2015. M33 Coimisiún na Scrúduithe Stáit State Examinations Commission LEAVING CERTIFICATE EXAMINATION, 2015 CHEMISTRY ORDINARY LEVEL TUESDAY, 16 JUNE AFTERNOON 2.00 TO 5.00 400 MARKS Answer eight questions More information TITRATION: STANDARDIZATION OF A BASE AND ANALYSIS OF STOMACH ANTACID TABLETS 2009, 1996, 1973 by David A. Katz. All rights reserved. Reproduction permitted for education use provided original copyright More information Assessment Chapter Test A Chapter: States of Matter In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question. 1. The kinetic-molecular More information 413 POWER PLANT ENGINEERING PART-A 1. Define Power. Power is the rate at which energy is used (or) Energy/time. 2. What are the types of fuels? Solid fuel Liquid fuel Gaseous fuel (Any one among the above More information Summer Holidays Questions Chapter 1 1) Barium hydroxide reacts with hydrochloric acid. The initial concentration of the 1 st solution its 0.1M and the volume is 100ml. The initial concentration of the More information ION EXCHANGE FOR DUMMIES An introduction Water Water is a liquid. Water is made of water molecules (formula H 2 O). All natural waters contain some foreign substances, usually in small amounts. The water More information Continuous process of sodium bicarbonate production by Solvay method Manual to experiment nr 10 Instructor: Dr Tomasz S. Pawłowski 1 Goal of the experiment The goal of the experiment is introduction of More information

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